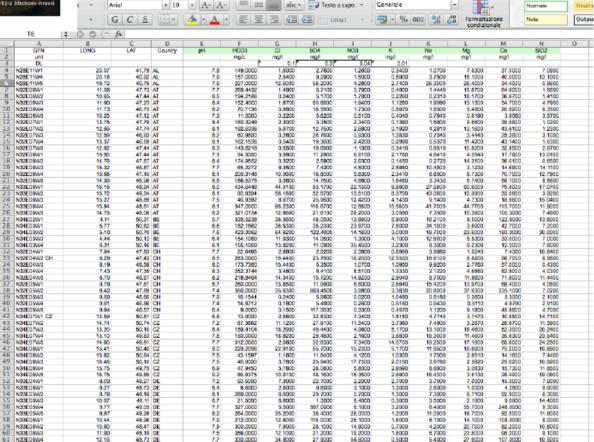
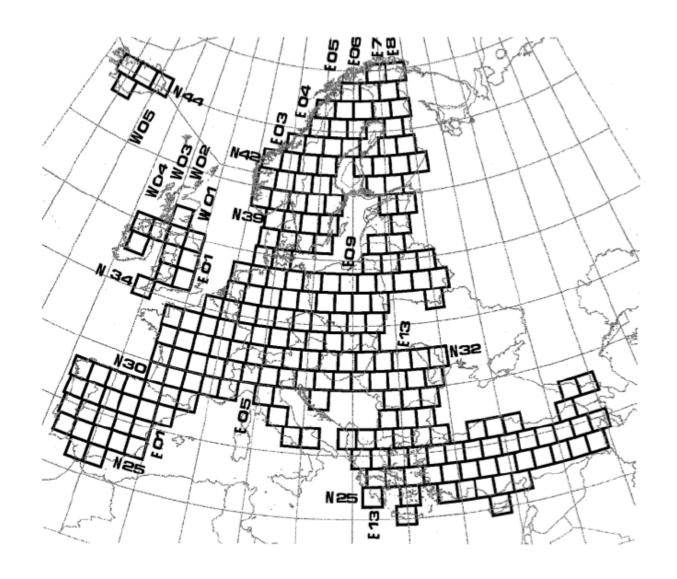


The FOREGS Geochemical
Baseline Mapping Programme's
main aim is to provide high
quality, multi-purpose
environmental geochemical
baseline data for Europe. The
need for this type of data was
justified by the first Working
Group on Regional Geochemical
Mapping immediately after the
Chernobyl accident in 1986, when
it was realised that a baseline for
radioactive and other polluting
elements could not be defined
(Bølviken et al. 1990, 1993, 1996).

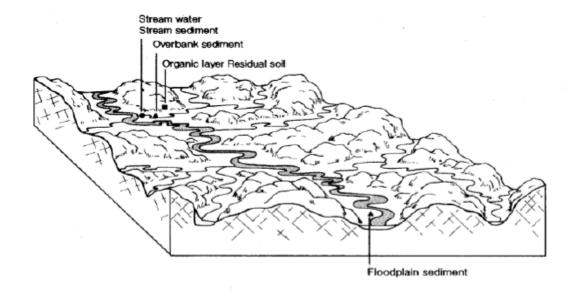


F 100% ■ (2)

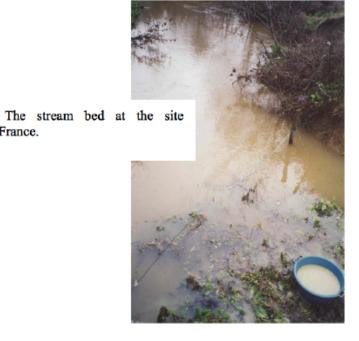
water0.xls



Global Terrestrial Network (GTN) cells in the FOREGS countries. (from Salminen, Tarvainen *et al.* 1998, Fig. 1, p.11). After this original plan, more cells were introduced to Greece, Italy and Spain to cover the coastal areas.



Block diagram showing possible sampling sites of GTN sampling media (from Salminen, Tarvainen et al. 1998, Fig.4, p.14; modified after Strahler 1969).



32E01W4, France.



Landscape about the water and stream sediment sampling site 32E01W4, France.

The IUGS/IAGC Global Geochemical Baselines Programme aims to establish a global geochemical reference baseline for >60 determinants in a range of media for environmental and other applications. The European contribution to the programme has been carried out by government institutions from 26 countries under the auspices of the former Forum of European Geological Surveys (FOREGS). This activity is now transferred to EuroGeoSurveys, the Association of the European Geological Surveys in which FOREGS is merged. The main objectives of this European survey were:

- to apply standardised methods of sampling, chemical analysis and data management to prepare a geochemical baseline across Europe;
 - 2) to use this reference network to level national baseline datasets.

Data on **geochemical baselines** are urgently needed in Europe, because environmental authorities in most countries are defining limits for contaminants in soils for different land use purposes. At the same time, the Commission of the European Union (EU) is preparing the Soil Protection Directive. As geochemists know, the natural concentrations of elements are different in the different constituents of overburden, and vary markedly between geologically disparate areas. State authorities, however, are not always aware of these significant natural variations, which should be taken into account in defining action limits. There are already examples of action limits that are lower than natural concentrations.



Figure 12. Figure 12. Filtration of a stream water sample (photo: Jari Višitšinen, GTK, from Salminen, Tarvainen et al., 1998, Fig. 6, p.17)



Figure 13. Figure 13. Wet sieving of a stream sediment sample in the UK (Photo: Fiona Fordyce, BGS from Salminen and Tarvainen et al. 1998, Fig. 7, p.21).

Table 2. The number of different sample types collected in each country.

Table 2. The han	Floodplain	Humus	Sub	Тор	Stream	Stream
	sediment		soil	soil	sediment	water
Albania	3		2	3	3	
Austria	19	12	15	18	20	20
Belgium	5	4	5	5	5	5
Croatia	12	12	13	14	13	10
Czech Republic	10	7	10	10	10	10
Denmark		1	5	5	5	5
Estonia	11	11	11	11	11	11
Finland	42	65	65	66	65	65
France	113	18	118	119	119	119
Germany	53	41	72	73	74	74
Greece	41		41	41	41	27
Hungary	14	3	14	14	14	10
Ireland	7	7	11	11	11	11
Italy	50	15	47	52	52	48
Latvia	7		8	7	7	7
Lithuania	13	5	15	1 6	16	14
Norway	47	58	58	58	58	58
Poland	56	29	56	56	56	56
Portugal	19		19	1 9	19	19
Slovak Republic	15	1	15	15	15	15
Slovenia	3	4	5	5	5	4
Spain	103		58	105	104	87
Sweden	47	51	51	51		51
Switzerland	10	9	9	1 0	10	10
The Netherlands		4	7	8	9	9
UK	51	28	60	60	60	60
Total	751	385	790	852	802	805

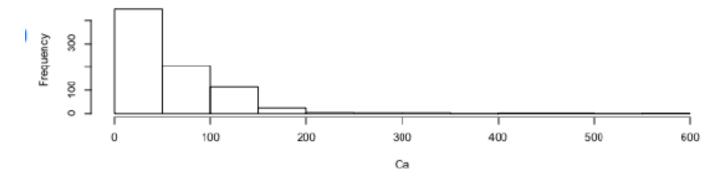
Analisi grafico-numerica con unità di concentrazione differenti

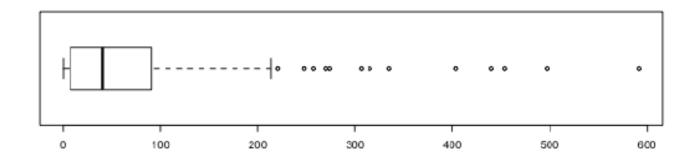
```
mmol/L = (mg/L) / (gram \ formula \ weight)
mmol/L = (ppm) \ . \ (density \ of \ sample) / (gram \ formula \ weight)
mmol/L = (meq/L) / \ (charge \ of \ ion)
mmol/l = molality \times density \times \frac{(weight \ solution - weight \ solutes)}{(weight \ soslution)} \times 100
(0.052 \ mmol/L \ Na^+) \times 1 = 0.052 \ meq/L
(1.8 \ mmol/L \ Ca^{2+}) \times 2 = 3.6 \ meq/L
(0.41 \ mmol/L \ SO_4^{2-}) \times 2 = -0.82 \ meq/L
```

Ions are electrically charged and the sums of positive and negative charges in a given water sample must balance. This condition is termede the **electroneutrality** of the solution. Since mmol/L represents the number of molecules, it should be multiplied by the charge of the ions to yield their total charge in meq/L.

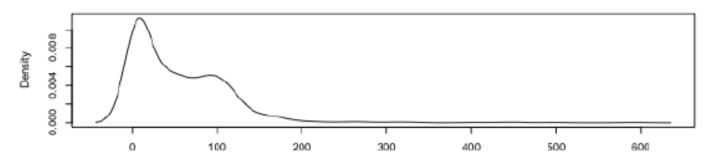
Scatter plots (covariance and correlation) of Ca²⁺ and HCO₃⁻ have to be investigate also in meq/L.

Histogram of Ca





density.default(x = Ca)



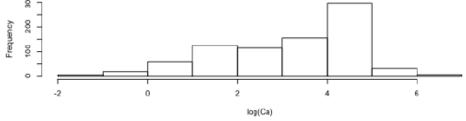
Ca²⁺

Sali ciclici = 0.1% Carbonati = 65% Silicati = 18% Evaporiti = 8%

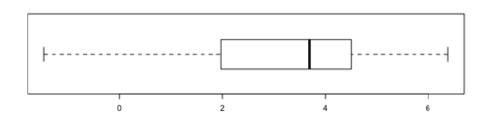
weathering

Inquinamento = 9%

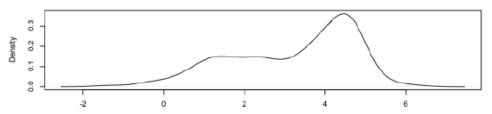




Histogram of log(Ca)







Ca (mg/L) stream water Europe

